

Thermochemistry Questions And Answers

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Thermochemistry Exam1 and Problem Solutions 1. Which ones of the following reactions are endothermic in other words ΔH is positive? I. $\text{H}_2\text{O}(l) + 10,5\text{kcal} \rightarrow \text{H}_2\text{O}(g)$ ΔH_1 II. $2\text{NH}_3 + 22\text{kcal}$

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Thermochemistry Thermochemistry and Energy and Temperature Thermochemistry is study of changes in energy (heat) associated with physical or chemical changes. Force = push $F = m a$ (mass x acceleration) force units: N (newton) = kg m s^{-2} Work = force x distance $W = F d$ energy units: J (joule) = $\text{kg m}^2 \text{s}^{-2}$

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An endothermic process the system _____ heat as the surroundings _____ heat. Q. Kinetic energy is the stored energy or energy of an object's position. Q. Thermochemistry is the study of _ changes during chemical and physical reactions. Q. Energy can be converted from one form to another.

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The answer is given per mole of reaction. $\Delta_f H^\circ$ (298 K) for SF_4 (g), H_2O (l), SO_2 (g) and HF (g) are -763 , -286 , -297 and -273 kJ mol $^{-1}$, respectively. What is the value of $\Delta_r H^\circ$ (298 K) for the hydrolysis shown above? For this process, $\Delta_f H^\circ$ (298 K) = -3953 kJ mol $^{-1}$.

~~Chapter 2: Thermochemistry~~

Thermochemistry practice problems 1) How can energy be transferred to or from a system? A) Energy can only be transferred as potential energy being converted to kinetic energy. B) Energy can be transferred only as heat. C) Energy can be transferred only as work. D) Energy can be transferred as heat and/or work.

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a) first eq \square flip; multiply by 3 \square 2 (this gives 3NO 2 as well as the 3NO which will be necessary to get one NO in the final answer) b) second eq \square divide by 2 (gives two nitric acid in the final answer) c) third eq \square flip (cancels 2NO as well as nitrogen) 2) Comment on the oxygens:

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