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unit in chemistry,

the mole, has a

special name,

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number, in honor of

the Italian scientist

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(1776-1856). The

commonly accepted

definition of

Avogadro ' s number

is the number of

atoms in exactly 12

g of the isotope ^{12}C

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C, and the quantity
itself is 6.02214199
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compare direct and indirect

measurement while applying proper measurement skills and significant digit rules. Lab Setup:

You will not need any materials for this lab. Instead, view the 1.05 Lab Introduction to get an overview. You

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will then use the lab photos to collect your data and complete the analysis.

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Analysis of Alum

Part II Determine

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the Water Lab 1

Hydration 8. Heat a crucible with cover over a burner flame until it is red hot.

Allow the crucible to cool, and then measure the total mass of crucible and cover. Handle the crucible with tongs or forceps to avoid getting fingerprints on it. 9.

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~~LabQuest 4~~

~~ANSWERS
PROCEDURE Part I~~

~~Determine the
Melting ...~~

Figure 1. PRE-LAB:

1. Set up the lab apparatus as directed by your teacher. See figure 1.
2. Turn on the computer. Double-click the Logger Pro icon. In the

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Logger Pro Lab 1

software select,

File, Open,

Experiment #5

“ Molar Volume ” in

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For each trial you

will use 5mL of 1M

hydrochloric acid

(CAUTION ...

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Molality Calculation

Molality Ethylene

glycol 0 mL water

5.0 mL 5 grams 0

grams mols 4.5 mL

4.5 grams 0.5 mL

0.45 grams mols kg

4.0 mL 4.0 grams

1.0 mL 0.90 grams

mols kg 3.5 mL 3.5

grams 1.5 mL 1.35

grams 2.0 mL 3.0

mL 3.0 grams 1.80

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grams mols 2.5 mL

2.5 grams kg 2.5

mL 2.25 grams

mols

~~Solved: Experiment~~

~~Freezing Point 2~~

~~Depression Adapted~~

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