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### Chapter 18 Chemical Equilibrium Answers

chemical equation for the reaction at equilibrium is written using double arrows to indicate the overall reversibility of the reaction.  $2\text{HgO}(s) \rightleftharpoons 2\text{Hg}(l) + \text{O}_2(g)$  Equilibrium, a Dynamic State Many chemical reactions are reversible under ordinary conditions of temperature and concentration. They will reach a state of equilibrium

### CHAPTER 18 Chemical Equilibrium

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$\text{H}_2(\text{g}) + \text{Br}_2(\text{g}) \rightleftharpoons 2\text{HBr}(\text{g})$   $K = 4.0 \times 10^{-2}$   $K = 3.5 \times 10^4$  Chapter 18 Chemical Equilibrium 18.2 Shifting Equilibrium The Effect of Change in Pressure: Ways to Change Pressure Add or remove a gaseous reactant or product Add an inert gas (one not involved in the reaction) An inert gas increases the total pressure but has no effect on the concentrations or partial pressures of the reactants or products The Effect of Change in Pressure Change the volume of the container When the volume of the ...

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SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_\_

Raising the temperature of any equilibrium system always (a) favors the forward reaction. (b) favors the

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reverse reaction. (c) favors the exothermic reaction.

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562 Chapter 18 Chemical Equilibrium Figure 18-3 If the only way in or out of San Francisco and Sausalito is across the Golden Gate Bridge, which joins the two cities, then the number of vehicles in the two cities will remain constant if the number of vehicles per hour crossing the bridge in one direction equals the number of vehi-

### Chapter 18: Chemical Equilibrium

5. Amongst the following chemical reactions, the irreversible reaction is: a)  $\text{AgNO}_3 + \text{NaCl} \rightarrow \text{AgCl} + \text{NaNO}_3$ . b)  $\text{H}_2 + \text{I}_2 \rightarrow \text{HI}$ . c)  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ . d)  $\text{O}_2 + 2\text{SO}_2 \rightarrow 2\text{SO}_3$ . ANS: The reaction is not reversible. 6. When the rate of forward reaction becomes equal to backward reaction, this state is termed as: a) Chemical equilibrium. b) Reversible state

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